

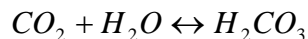
THE EFFECT OF ACID RAIN ON OCEAN AND LAKE WATER

LAB ACID 18

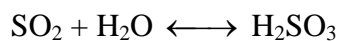
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INTRODUCTION

Acid rain is a topic of much concern in today's world. As carbon dioxide gas, CO_2 , dissolves in water droplets of unpolluted air, the following reaction occurs:



H_2CO_3 is a weak acid that causes the rain from unpolluted air to be slightly acidic. This source of "acid rain" is not usually considered to be a pollutant, since it is natural and usually does not alter the pH of rain water very much. Oxides of sulfur dissolve in water droplets to cause more serious problems. Sulfur dioxide dissolves to produce sulfurous acid, H_2SO_3 , by the equation:



This source of sulfur dioxide can occur naturally, as from volcanic gases. More often, however, sulfur dioxide is considered a pollutant, since it is a by-product of fossil-fuel combustion.

The acidity of a solution can be expressed using the pH scale, which ranges from 0 to 14. Solutions with a pH above 7 are basic, solutions with pH below 7 are acidic, and a neutral solution has a pH of 7. In Part I of this experiment, you will study how the pH of water changes when CO_2 is dissolved in water. In Part II, you will study the effect sulfuric acid has on the pH of different water types.

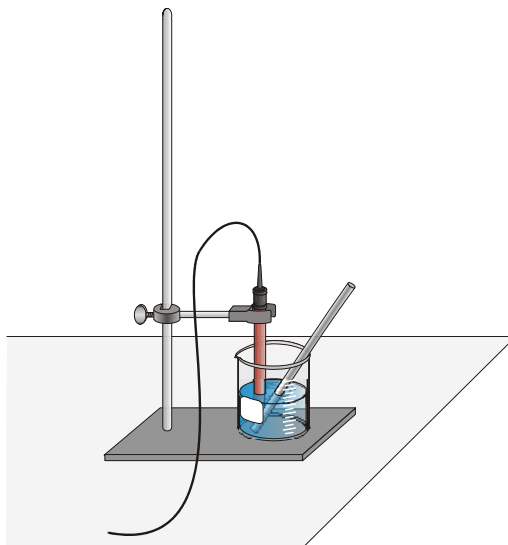


Figure 1

MATERIALS

LabPro or CBL 2 interface	dilute H ₂ SO ₄
TI Graphing Calculator	ring stand and utility clamp
DataMate program	straw
Vernier pH Sensor	wash bottle with distilled water
Graphical Analysis or graph paper	water from a lake
250-mL and 100-mL beaker	water from the ocean (optional)
buffer solution	

PROCEDURE

Part I CO₂ and Water

As carbon dioxide gas, CO₂, dissolves into water droplets suspended in the atmosphere, it changes the rainwater's pH. Here you will test to see if CO₂ will affect the pH of distilled water. The source of CO₂ will be your breath.

1. Obtain and wear goggles.
2. Before each use of the pH Sensor, you need to rinse the tip of the sensor thoroughly with distilled water. To do this, hold the pH Sensor above a rinse beaker and use the rinse bottle to thoroughly rinse the sensor tip.

Important: Do not let the pH Sensor dry out. Keep it in a 250-mL beaker with about 100 mL of tap water when not in use. The tip of the sensor is made of glass—it is very fragile. Handle with care!

3. Plug the pH Sensor into Channel 1 of the LabPro or CBL 2 interface. Use the link cable to connect the TI Graphing Calculator to the interface. Firmly press in the cable ends.
4. Turn on the calculator and start the DATAMATE program. Press **CLEAR** to reset the program.
5. Set up the calculator and interface for a pH System.
 - a. If the calculator displays PH in CH 1, proceed directly to Step 6. If it does not, continue with this step to set up your sensor manually.
 - b. Select SETUP from the main screen.
 - c. Press **ENTER** to select CH 1.
 - d. Select PH from the SELECT SENSOR menu.
 - e. Select OK to return to the main screen.
6. Wash a 250-mL beaker with tap water and dry it with a paper towel. Note: All glassware must be clean in this experiment! Put 100 mL of distilled water into this clean beaker. Using a ring-stand and utility clamp, secure the pH Sensor in the beaker as shown in Figure 1. The tip of the sensor should be down in the distilled water. Swirl the water around the sensor tip briefly. Place a clean straw into the distilled water.
7. Record the pH value displayed on the screen as the initial pH (time 0) in Table 1.
8. Begin data collection by selecting START on the calculator. Take a deep breath, then gently blow your breath into the distilled water for as long as possible.

- When data collection is finished, a graph of pH vs. time will be displayed on the calculator screen. As you move the cursor right or left, the time (X) and pH (Y) values of each data point are displayed below the graph. In Table 1, record the pH value at 10, 20, 30, 40, 50, 60, 70, 80, and 90 seconds.

Part II Simulating Acid Rain Using H₂SO₄

- Empty the distilled water from the beaker and rinse the pH Sensor thoroughly with distilled water from a rinse bottle. Place the pH Sensor in the storage beaker with tap water.
- Wash and dry the 250-mL beaker. Fill the beaker with a 50-mL portion of distilled water. Set up the beaker and pH Sensor as shown in Figure 1.
- Set up the data-collection mode.
 - Select SETUP from the main screen.
 - To select MODE, press (the up arrow key) once and press .
 - Select EVENTS WITH ENTRY from the SELECT MODE menu.
 - Select OK to return to the main screen.
- Select START to begin taking pH measurements. Monitor the pH readings displayed on the calculator screen. When the readings are stable, press on the calculator.
- The calculator will prompt you for the number of drops of acid you added to the beaker. Type "0" and press .
- Add 1 drop of acid to the beaker. Stir the solution thoroughly after addition. **CAUTION:** *Handle the sulfuric acid with care. It can cause painful burns if it comes in contact with the skin. Avoid spilling it on your skin or clothing.*
- When the readings are stable, press on the calculator. Enter the total number of drops of acid added to the water in the beaker and press .
- Repeat Steps 15 – 16, adding 1 drop each time until a total of 10 drops has been added.
- On the calculator, press to stop data collection. A graph will be displayed with pH along the y-axis and the number of drops added along the x-axis. Move the flashing cursor from point to point and record the pH values in Table 2. When finished, press to return to the main screen.

H₂SO₄ and Water from the Ocean

- Repeat Steps 13 – 18 using "Ocean Water" in place of distilled water.

H₂SO₄ and Lake Water

- Repeat Steps 13 – 18 using "Lake Water" in place of distilled water.

H₂SO₄ and Buffer Solution

- Repeat Steps 13 – 18 using "Buffer Solution" in place of distilled water.
- Using Graphical Analysis, or by hand, make a plot of pH vs. drops of acid added. Place all four graphs (distilled, ocean, lake, and buffered water) on the same plot.

DATA TABLE

Part I

Table 1: Adding CO ₂ from your breath to water											
Time (s)	0	10	20	30	40	50	60	70	80	90	ΔpH
pH											

Part II

Table 2				
Drops	pH of this water type			
	distilled	ocean	lake	buffer
0				
1				
2				
3				
4				
5				
6				
7				
8				
9				
10				
ΔpH				

QUESTIONS

1. Calculate the change in pH (ΔpH) for the water in Part I. Subtract the final pH from the initial pH. What conclusion can you make about your breath?
2. Why does the pH change rapidly at first, and remain stable after a time?
3. Calculate the change in pH (ΔpH) for each of the Part II trials. Subtract the final pH from the initial pH.

4. Compare the ΔpH values. Which test gave the largest pH change? Which test gave the smallest pH change?
5. Water from the ocean is said to be “naturally buffered.” From the result of this experiment, what does this mean?
6. How does water from the ocean become buffered?
7. Many aquatic life forms can only survive in water with a narrow range of pH values. In which body of water—lakes or oceans—would living things be more threatened by acid rain? Explain.
8. There are numerous coal-burning electric power plants along the Ohio River in Southern Indiana, where the river and lake waters are naturally buffered. However, air pollution produced there is more harmful to water life in Upstate New York, where the river and lake waters are NOT buffered like those in Southern Indiana. A similar situation exists in Europe where air pollutants from highly industrialized Germany are more harmful to Scandinavian water life than to water life in Germany. Use the results of this experiment to explain these situations.
9. Summarize your conclusions of this laboratory experiment. Use your data to answer the purposes of this experiment.

EXTENSION

1. Test hard and soft water in the same way you tested lake and ocean water. How do they compare?
2. Do library research to get more information on the effects of acid rain on streams and lakes.
3. Do library research and prepare a report on “naturally buffered” streams and lakes.